

IGCSE Chemistry 4335/2H

Mark Scheme (Results)

November 2008

IGCSE

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The following acronyms are used

owtte or words to that effect

ecf error carried forward

dop dependent on previous

nwn no working necessary

Question Number	Correct Answer	Notes	Mark
1 (a) (i)	hydrogen peroxide → water + oxygen		(1)

Question Number	Correct Answer	Notes	Mark
1 (a) (ii)	catalyst		(1)

Question Number	Correct Answer	Notes	Mark
1 (b)	over water / displacement of air with downward delivery / upward displacement of air. <i>Could be shown on a diagram.</i>	Accept “through water”.	(1)

Question Number	Correct Answer	Notes	Mark
1 (c)	relights a glowing splint	Reject “glows more brightly”	(1)

Question Number	Correct Answer	Notes	Mark
1 (d) (i)	Red (ignore pale/dark), crimson / scarlet	Reject references to orange / yellow / pink	(1)

Question Number	Correct Answer	Notes	Mark
1 (d) (ii)	electron transfer from lithium to oxygen Li atoms each lose one electron and O atom gains two electrons	Covalent / sharing scores zero	1 1 1 (3)

Question Number	Correct Answer	Notes	Mark
1 (d) (iii)	Li ⁺ O ²⁻	Both correct but reversed scores 1	1 1 (2)

(Total 10 marks)

Question Number	Correct Answer	Notes	Mark
2 (a)	Brown / red brown (reject “light”, accept “dark”)	Reject red alone or reference to orange	1
	Grey (reject “light”, accept “dark”)/ black	Reject purple or violet	1 (2)

Question Number	Correct Answer	Notes	Mark
2 (b) (i)	diffusion		(1)

Question Number	Correct Answer	Notes	Mark
2 (b) (ii)	Br ₂ (l) → Br ₂ (g) <i>Reactants = 1, products = 1</i>		(2)

Question Number	Correct Answer	Notes	Mark
2 (b) (iii)	moving (faster) further apart owtte		1
			1 (2)

Question Number	Correct Answer	Notes	Mark
2 (c) (i)	bromine + hydrogen → hydrogen bromide	Ignore “gas”	(1)

Question Number	Correct Answer	Notes	Mark
2 (c) (ii)	hydrobromic (acid)		(1)

Question Number	Correct Answer	Notes	Mark
2 (d) (i)	melt/molten/fused (lead (II) bromide)	Allow “dissolve in water” not just heat or add water	(1)

Question Number	Correct Answer	Notes	Mark
2 (d) (ii)	(A) electrons	If ions named in wrong order, award 1 mark for both. If “ions” omitted twice, max 1 for B and C. If B and C “positive ions” and “negative ions” - 1 mark.	1
	(B) lead(II) ions / Pb ²⁺		1
	(C) bromide ions / Br ⁻		1
			(3)

Question Number	Correct Answer	Notes	Mark
2 (e)	Gain of electrons (by Pb^{2+})		(1)

(Total 14 marks)

Question Number	Correct Answer	Notes	Mark
4 (a) (i)	$C_2H_4 + H_2O \rightarrow CH_3CH_2OH$	Accept C_2H_5OH	(1)

Question Number	Correct Answer	Notes	Mark
4 (a) (ii)	(concentrated) phosphoric acid	Reject "dilute"	(1)

Question Number	Correct Answer	Notes	Mark
4 (b)	sugar / sucrose / glucose yeast two from dissolve in water absence of air temperature in range 20 - 40 °C <i>Any two conditions for 1 mark each</i>	Additional wrong reagents negate sugar mark	1
			1
		Reject "heat" accept "warm"	2
			(4)

Question Number	Correct Answer	Notes	Mark
4 (c) (i)	$CH_3CH_2OH(l) + CH_3COOH(l) \rightarrow CH_3COOCH_2CH_3(l) + H_2O(l)$ Reactants = 1, products = 1, state symbols (dependent on correct formulae) = 1	Accept C_2H_5 in place of CH_2CH_3	(3)

Question Number	Correct Answer	Notes	Mark
4 (c) (ii)	pleasant/fruity/glue smell / oily drops	Not just "smell"	(1)

(Total 10 marks)

Question Number	Correct Answer	Notes	Mark
5 (a) (i)	number of electrons in outer shell is same as group OR number of shells with electrons in is same as period		(1)

Question Number	Correct Answer	Notes	Mark
5 (a) (ii)	2.8.8.2	Accept any punctuation	(1)

Question Number	Correct Answer	Notes	Mark
5 (b)	ATOMS with (If atoms omitted, max 1) same atomic number/same number of protons/same element(1) different numbers of neutrons/mass number (1)	Ignore same electrons	(2)

Question Number	Correct Answer					Notes	Mark
5 (c) (i)	Number of neutrons	Number of protons	Atomic number of isotope	Mass number of isotope	Percentage isotope in the element		(5)
	12 (1)	12(1)	12	24	79		
	13	12	12	25(1)	10(1)		
	14	12	12(1)	26	11		

Question Number	Correct Answer	Notes	Mark
5 (c) (ii)	Magnesium/Mg		(1)

Question Number	Correct Answer	Notes	Mark
5 (c) (iii)	cq on percentages in table. If use only two isotopes max 1. evidence of multiplication of mass numbers by percentages <i>correct answer answer to 3 sig figs.</i> $24.3 = 3$ $24.32 = 2$	If divide by 10 or 1000 rather than 100, max 1 First step nonsense = 0	1
			1
			1
			(3)

Question Number	Correct Answer	Notes	Mark
5 (c) (iv)	Effervescence/ bubbles/ same/no difference same electronic configuration / same element / same number of electrons / number of neutrons has no effect		1
			1
			(2)

(Total 15 marks)

Question Number	Correct Answer	Notes	Mark
6 (a) (i)	electrons able to move/can flow/mobile (dependent on first mark)	Ignore "delocalised"	1 1 (2)

Question Number	Correct Answer	Notes	Mark
6 (a) (ii)	layers/rows/lines (of ions /atoms) not electrons slide/slip / move over each other (dependent on first mark) <i>do not accept answers based on atoms moving into gaps or electrons acting as ball bearings.</i>		1 1 (2)

Question Number	Correct Answer	Notes	Mark
6 (b)	left hand electrode labelled (pure) copper right hand electrode labelled impure copper electrolyte labelled as any soluble copper salt (solution)	Accept cathode Accept anode	1 1 1 (3)

Question Number	Correct Answer	Notes	Mark
6 (c) (i)	solution has lower melting point/melting point of aluminium oxide is too high. <i>allow lowers mp of aluminium oxide.</i>		(1)

Question Number	Correct Answer	Notes	Mark
6 (c) (ii)	Carbon (accept graphite)		(1)

Question Number	Correct Answer	Notes	Mark
6 (d)	Copper: electrical wires / coins / water pipes / allow pans / Associated property (conductor must be qualified). Aluminium: overhead cables/ specified transport/ pans / cooking foil / drink cans Associated property (conductor must be qualified).	Reject coins	1 1 1 1 (4)

Question Number	Correct Answer	Notes	Mark
6 (e)	either: electrolysis (1) more reactive than C/can not be reduced by C/similar reactivity to Al/Al is extracted by electrolysis. (1) or: react with a NAMED more reactive metal (1) Ti less reactive than metal used/metal used more reactive than Ti/ metal will displace Ti. (1)		(2)

(Total 15 marks)

Question Number	Correct Answer	Notes	Mark
7 (a)	$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$ reagents = 1 products = 1 balanced = 1 (dependent on first two marks given)		(3)

Question Number	Correct Answer	Notes	Mark
7 (b)	exothermic/gives out (heat) energy		(1)

Question Number	Correct Answer	Notes	Mark
7 (c)	<ul style="list-style-type: none"> • make chlorides into solutions/add water • green ppt • brown ppt • correct linking of at least one observation to a cation 		(4)

(Total 8 marks)

Question Number	Correct Answer	Notes	Mark
8 (a) (i)	contain oxygen/contains an element other than C and H		(1)

Question Number	Correct Answer	Notes	Mark
8 (a) (ii)	CH ₃ / H ₃ C		(1)

Question Number	Correct Answer	Notes	Mark
8 (a) (iii)	any TWO from <ul style="list-style-type: none"> • same general formula • members differ by CH₂ • same/similar chemical reactions /same functional group • gradation in physical properties 	Accept trend in stated property	(2)

Question Number	Correct Answer	Notes	Mark
8 (a) (iv)	6 C-H bond pairs and one C-C bond pair. No other outer electrons Not ethane = 0	Ignore inner electrons	1 1 (2)

Question Number	Correct Answer	Notes	Mark
8 (a) (v)	poly(propene)/polypropene/polypropylene		(1)

Question Number	Correct Answer	Notes	Mark
8 (a) (vi)	1 correct repeat unit shown with continuation bonds (dependent on correct structure)		1 1 (2)

Question Number	Correct Answer	Notes	Mark
8 (a) (vii)	E has double bond/unsaturated polymer no double bond/saturated		1 1 (2)

Question Number	Correct Answer	Notes	Mark
8 (b)	three correct structures from: but-1-ene but-2-ene methylpropene cyclobutane methylcyclopropane	Penalise CH ₃ or CH ₂ once Penalise sticks once	(3)

(Total 14 marks)

Question Number	Correct Answer	Notes	Mark
9 (a) (i)	= 64	Ignore units	(1)

Question Number	Correct Answer	Notes	Mark
9 (a) (ii)	2 (cq) (128/a(i))		(1)

Question Number	Correct Answer	Notes	Mark
9 (a) (iii)	6 (cq) (a(ii) x 3)		(1)

Question Number	Correct Answer	Notes	Mark
9 (a) (iv)	6x12 =72 cq (a(iii) x 12)	Units not required, incorrect units lose mark.	(1)

Question Number	Correct Answer	Notes	Mark
9 (b) (i)	Ca(OH) ₂		(1)

Question Number	Correct Answer	Notes	Mark
9 (b) (ii)	water/H ₂ O carbon dioxide/CO ₂		1 1 (2)

Question Number	Correct Answer	Notes	Mark
9 (c) (i)	correct bonds/numbers identified (2 x 412 + 1 x 837 + 2 x 431) 2523 <i>correct answer = 2. cq on 1 error in calculation for max 1</i>		1 1 (2)

Question Number	Correct Answer	Notes	Mark
9 (c) (ii)	correct bonds/numbers identified (4 x 412 + 1 x 348 + 2 x 338) 2672 <i>correct answer = 2. cq on 1 error in calculation for max 1</i>		1 1 (2)

Question Number	Correct Answer	Notes	Mark
9 (c) (iii)	$Cq c(i) - c(ii) - 149$		(1)

(Total 12 marks)

Question Number	Correct Answer	Notes	Mark
10 (a)	giant / macromolecular	Reject ionic	(1)

Question Number	Correct Answer	Notes	Mark
10 (b)	<ul style="list-style-type: none"> break covalent bonds (between atoms) covalent bonds strong need lots of energy to overcome/break 	If ionic / hydrogen bonds / vdw forces / delocalised electrons / molecules = 0	(3)

Question Number	Correct Answer	Notes	Mark
10 (c)	<ul style="list-style-type: none"> weak forces between layer slide/slip 		(2)

Question Number	Correct Answer	Notes	Mark
10 (d) (i)	<ul style="list-style-type: none"> weak forces between molecules little energy to overcome no (covalent) bonds broken / in diamond (covalent) bonds broken 		(3)

Question Number	Correct Answer	Notes	Mark
10 (d) (ii)	<p>if yes:</p> <p>any two from</p> <ul style="list-style-type: none"> (molecules) round/balls/football shaped weak forces between molecules roll <p>if no:</p> <ul style="list-style-type: none"> (strong) covalent bonds hold atoms in place/need lots of energy to break (dependent on M1) 		(2)

(Total 11 marks)