Instructions

- Use black ink or ball-point pen.
- Fill in the boxes at the top of this page with your name, centre number and candidate number.
- Answer all questions.
- Answer the questions in the spaces provided—there may be more space than you need.

Information

- The total mark for this paper is 80.
- The marks for each question are shown in brackets—use this as a guide as to how much time to spend on each question.
- Questions labelled with an asterisk (*) are ones where the quality of your written communication will be assessed— you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.
### SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 20 minutes on this section. For each question, select one answer from A to D and put a cross in the box \( \checkmark \). If you change your mind, put a line through the box \( \times \) and then mark your new answer with a cross \( \checkmark \).

1. What are the shapes of the BF₃ and PH₃ molecules?

<table>
<thead>
<tr>
<th></th>
<th>BF₃</th>
<th>PH₃</th>
</tr>
</thead>
<tbody>
<tr>
<td>A</td>
<td>pyramidal</td>
<td>pyramidal</td>
</tr>
<tr>
<td>B</td>
<td>pyramidal</td>
<td>trigonal planar</td>
</tr>
<tr>
<td>C</td>
<td>trigonal planar</td>
<td>pyramidal</td>
</tr>
<tr>
<td>D</td>
<td>trigonal planar</td>
<td>trigonal planar</td>
</tr>
</tbody>
</table>

(Total for Question 1 = 1 mark)

2. What are the C—C—C bond angles in diamond and graphite?

<table>
<thead>
<tr>
<th></th>
<th>Diamond</th>
<th>Graphite</th>
</tr>
</thead>
<tbody>
<tr>
<td>A</td>
<td>109.5°</td>
<td>109.5°</td>
</tr>
<tr>
<td>B</td>
<td>109.5°</td>
<td>120°</td>
</tr>
<tr>
<td>C</td>
<td>120°</td>
<td>109.5°</td>
</tr>
<tr>
<td>D</td>
<td>120°</td>
<td>120°</td>
</tr>
</tbody>
</table>

(Total for Question 2 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
3. Which describes the polarity of the C—Cl bond and the polarity of the molecule trichloromethane, CHCl₃?

<table>
<thead>
<tr>
<th>Polarity of C—Cl bond</th>
<th>Polarity of molecule</th>
</tr>
</thead>
<tbody>
<tr>
<td>□ A non-polar</td>
<td>non-polar</td>
</tr>
<tr>
<td>□ B non-polar</td>
<td>polar</td>
</tr>
<tr>
<td>□ C polar</td>
<td>non-polar</td>
</tr>
<tr>
<td>□ D polar</td>
<td>polar</td>
</tr>
</tbody>
</table>

(Total for Question 3 = 1 mark)

4. Which isomer, with the formula C₇H₁₆, will have the **lowest** boiling temperature?

- □ A CH₃CH₂CH₂CH₂CH₂CH₂CH₃
- □ B (CH₃)₂CHCH₂CH₂CH₂CH₃
- □ C CH₃CH₂C(CH₃)₂CH₂CH₃
- □ D (CH₃)₂CHC(CH₃)₃

(Total for Question 4 = 1 mark)

5. Which is a disproportionation reaction?

- □ A CaCO₃ → CaO + CO₂
- □ B 2H₂O₂ → 2H₂O + O₂
- □ C 2H₂S + 3O₂ → 2SO₂ + 2H₂O
- □ D Mg(OH)₂ → MgO + H₂O

(Total for Question 5 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
6. Which shows the trend in solubility of the hydroxides and sulfates of the Group 2 elements going up the group from barium to magnesium?

<table>
<thead>
<tr>
<th></th>
<th>Solubility of Group 2 hydroxides</th>
<th>Solubility of Group 2 sulfates</th>
</tr>
</thead>
<tbody>
<tr>
<td>A</td>
<td>decreases</td>
<td>decreases</td>
</tr>
<tr>
<td>B</td>
<td>decreases</td>
<td>increases</td>
</tr>
<tr>
<td>C</td>
<td>increases</td>
<td>decreases</td>
</tr>
<tr>
<td>D</td>
<td>increases</td>
<td>increases</td>
</tr>
</tbody>
</table>

(Total for Question 6 = 1 mark)

7. Which diagram shows a Maxwell-Boltzmann distribution of molecular energies?

- A
  - Fraction of molecules with energy, $E$ vs. energy, $E$

- B
  - Fraction of molecules with energy, $E$ vs. energy, $E$

- C
  - Fraction of molecules with energy, $E$ vs. energy, $E$

- D
  - Fraction of molecules with energy, $E$ vs. energy, $E$

(Total for Question 7 = 1 mark)
8 The rate of the reaction between sodium thiosulfate solution and dilute hydrochloric acid increases as the concentration of sodium thiosulfate increases.

Which of these occurs when the concentration of the sodium thiosulfate solution increases at constant temperature?

<table>
<thead>
<tr>
<th>Activation energy</th>
<th>Particles</th>
</tr>
</thead>
<tbody>
<tr>
<td>□ A</td>
<td>decreases</td>
</tr>
<tr>
<td></td>
<td>collide more frequently</td>
</tr>
<tr>
<td>□ B</td>
<td>decreases</td>
</tr>
<tr>
<td></td>
<td>collide with more energy</td>
</tr>
<tr>
<td>□ C</td>
<td>stays the same</td>
</tr>
<tr>
<td></td>
<td>collide more frequently</td>
</tr>
<tr>
<td>□ D</td>
<td>stays the same</td>
</tr>
<tr>
<td></td>
<td>collide with more energy</td>
</tr>
</tbody>
</table>

(Total for Question 8 = 1 mark)

9 Consider the following exothermic reaction.

\[ A(s) + B(aq) \rightarrow C(aq) + D(g) \]

If the mass of A, and the volume and concentration of the solution of B are constant, which of these changes in conditions will result in the fastest initial rate?

<table>
<thead>
<tr>
<th>Size of solid particles of A</th>
<th>Temperature</th>
</tr>
</thead>
<tbody>
<tr>
<td>□ A</td>
<td>doubled</td>
</tr>
<tr>
<td></td>
<td>decreased by 10°C</td>
</tr>
<tr>
<td>□ B</td>
<td>doubled</td>
</tr>
<tr>
<td></td>
<td>increased by 10°C</td>
</tr>
<tr>
<td>□ C</td>
<td>halved</td>
</tr>
<tr>
<td></td>
<td>decreased by 10°C</td>
</tr>
<tr>
<td>□ D</td>
<td>halved</td>
</tr>
<tr>
<td></td>
<td>increased by 10°C</td>
</tr>
</tbody>
</table>

(Total for Question 9 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
10 The reaction profile for an exothermic catalysed reaction is shown.

![Reaction Profile Diagram]

Which arrow represents the activation energy for this reaction?

- [ ] A
- [ ] B
- [ ] C
- [ ] D

(Total for Question 10 = 1 mark)

11 An aqueous solution contains dichromate(VI) ions, \( \text{Cr}_2\text{O}_7^{2-} \), and chromate(VI) ions, \( \text{CrO}_4^{2-} \), in equilibrium. This solution is a pale orange colour.

\[
\text{Cr}_2\text{O}_7^{2-} (\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons 2\text{CrO}_4^{2-} (\text{aq}) + 2\text{H}^+ (\text{aq})
\]

orange \( \rightleftharpoons \) yellow

What would be seen when a few drops of concentrated sodium hydroxide solution are added to the equilibrium mixture?

- [ ] A No visible change.
- [ ] B The mixture turns green.
- [ ] C The mixture turns a deeper orange.
- [ ] D The mixture turns yellow.

(Total for Question 11 = 1 mark)
12 The following system was allowed to reach equilibrium at 450°C.

\[
2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g}) \quad \Delta H = -197 \text{ kJ mol}^{-1}
\]

How would a decrease in pressure and an increase in temperature affect the equilibrium position?

<table>
<thead>
<tr>
<th>Shift in equilibrium position with a decrease in pressure</th>
<th>Shift in equilibrium position with an increase in temperature</th>
</tr>
</thead>
<tbody>
<tr>
<td>□ A left</td>
<td>left</td>
</tr>
<tr>
<td>□ B left</td>
<td>right</td>
</tr>
<tr>
<td>□ C right</td>
<td>left</td>
</tr>
<tr>
<td>□ D right</td>
<td>right</td>
</tr>
</tbody>
</table>

(Total for Question 12 = 1 mark)

13 What is the empirical formula of a bromoalkane containing, by mass, 22.0% carbon, 4.6% hydrogen and 73.4% bromine?

(Relative atomic masses: C = 12, H = 1, Br = 80)

□ A C\textsubscript{3}H\textsubscript{7}Br

□ B C\textsubscript{2}H\textsubscript{5}Br

□ C C\textsubscript{2}H\textsubscript{3}Br

□ D CH\textsubscript{3}Br

(Total for Question 13 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
14 Four isomers with the formula C₄H₉Br are shown.

<p>| | |</p>
<table>
<thead>
<tr>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>P</td>
<td>Q</td>
</tr>
<tr>
<td>Br</td>
<td>Br</td>
</tr>
<tr>
<td>R</td>
<td>S</td>
</tr>
</tbody>
</table>

Which of the isomers are primary halogenoalkanes?

- [ ] A P and R
- [ ] B P and S
- [ ] C Q and R
- [ ] D Q only

(Total for Question 14 = 1 mark)

15 How many different alkenes could be formed when 2-iodopentane, CH₃CHICH₂CH₂CH₃, reacts with alcoholic potassium hydroxide?

- [ ] A 1
- [ ] B 2
- [ ] C 3
- [ ] D 4

(Total for Question 15 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
16 1-bromobutane can be made from butan-1-ol.

\[ \text{CH}_3\text{CH}_2\text{CH}_2\text{OH} + \text{HBr} \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{Br} + \text{H}_2\text{O} \]

What mass of 1-bromobutane is formed from 3.7 g of butan-1-ol if the yield is 56%?

(Relative molecular masses: butan-1-ol = 74, 1-bromobutane = 137)

- A 3.84 g
- B 6.85 g
- C 12.23 g
- D 76.72 g

(Total for Question 16 = 1 mark)

17 Which of these molecules does not absorb infrared radiation?

- A carbon monoxide
- B carbon dioxide
- C oxygen
- D water

(Total for Question 17 = 1 mark)

18 Glucose is fermented to produce ethanol.

\[ \text{C}_6\text{H}_12\text{O}_6 \rightarrow 2\text{C}_2\text{H}_5\text{OH} + 2\text{CO}_2 \]

What is the atom economy, by mass, for the production of ethanol in this reaction?

(Relative molecular masses: \text{C}_6\text{H}_12\text{O}_6 = 180, \text{C}_2\text{H}_5\text{OH} = 46, \text{CO}_2 = 44)

- A 25.6%
- B 48.9%
- C 50.0%
- D 51.1%

(Total for Question 18 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
19 The mass spectrum of propan-1-ol is shown.

Which peak represents the molecular ion for propan-1-ol containing a carbon-13 isotope?

- [ ] A
- [ ] B
- [ ] C
- [ ] D

(Total for Question 19 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.
20 Compounds containing oxygen are sometimes added to hydrocarbon fuels to reduce incomplete combustion and improve engine performance. Which contains the greatest number of oxygen atoms?

(Relative molecular masses: CH₃OH = 32, C₂H₅OH = 46, CH₂OHCH₂OH = 62, C₄H₉OH = 74)

☐ A  8.0 g of methanol, CH₃OH
☐ B  9.2 g of ethanol, C₂H₅OH
☐ C  6.2 g of ethane-1,2-diol, CH₂OHCH₂OH
☐ D  7.4 g of butan-1-ol, C₄H₉OH

(Total for Question 20 = 1 mark)
SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

21 This question is about the carbonates and nitrates of elements in Group 1 and Group 2 of the Periodic Table.

(a) Many of the metal ions of Group 1 and Group 2 can be identified using flame tests.

(i) State the colour given to a flame by barium nitrate.

(ii) Explain the origin of the flame colour.

(b) Sodium nitrate and magnesium nitrate decompose when they are heated.

Write equations to show the thermal decomposition of each of these nitrates. State symbols are not required.

(i) Sodium nitrate

(ii) Magnesium nitrate
(c) Magnesium carbonate decomposes readily when heated but sodium carbonate does not.

Explain this observation by including reference to the charge and size of the cations. (4)
(d) Hydrated sodium carbonate has the formula $\text{Na}_2\text{CO}_3\cdot x\text{H}_2\text{O}$.

A student determined the value of $x$ in the formula of a sample of hydrated sodium carbonate. The following procedure was used.

- Use 2.50 g of hydrated sodium carbonate to prepare 250 cm$^3$ of solution.
- Use a pipette to transfer 25.0 cm$^3$ of the sodium carbonate solution to a conical flask.
- Add a few drops of methyl orange indicator to the conical flask.
- Titrate the solution with 0.105 mol dm$^{-3}$ hydrochloric acid until concordant results are obtained.

The student’s mean titre was 16.65 cm$^3$.

The equation for the reaction is

$$\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$$

*(i) Calculate the amount, in moles, of sodium carbonate, $\text{Na}_2\text{CO}_3$, in the 250 cm$^3$ of solution in the volumetric flask.

(amount $\text{Na}_2\text{CO}_3$ in 250 cm$^3$) = ............................................. mol
(ii) Calculate the molar mass of Na$_2$CO$_3$.xH$_2$O and hence the value of $x$. 

(2)
(iii) Another student carried out the same experiment but obtained a different answer. The method this student used for preparing the sodium carbonate solution is shown.

I weighed 2.50 g of hydrated sodium carbonate in a weighing bottle and then tipped the solid into a 250 cm³ volumetric flask.
I dissolved the solid in a small amount of distilled water and then added distilled water up to the mark.
I then carried out a series of titrations.

Identify two errors that the student made in preparing this solution and explain the effect these errors will have on the titration volumes.

Error 1
........................................................................................................................................................................

Effect on the titration volumes
........................................................................................................................................................................
........................................................................................................................................................................
........................................................................................................................................................................

Error 2
........................................................................................................................................................................

Effect on the titration volumes
........................................................................................................................................................................
........................................................................................................................................................................
........................................................................................................................................................................

(Total for Question 21 = 19 marks)
22 This question concerns the halogens and some of their compounds.

(a) A halogen dissolves in water to form a yellow solution, and in cyclohexane to form a purple solution.

Name the halogen.

(b) Oxygen difluoride, OF₂, is produced in the reaction between fluorine and cold, dilute sodium hydroxide solution.

\[ 2F_2 + 2OH^- \rightarrow OF_2 + 2F^- + H_2O \]

Give the oxidation numbers of fluorine and oxygen in all of the species in the equation above and use them to explain why this is a redox reaction.
(c) Chlorine oxidises thiosulphate ions, $S_2O_3^{2-}$, to sulfate(VI) ions.

The ionic half-equations for the reaction are

$$\text{Cl}_2 + 2e^- \rightarrow 2\text{Cl}^-$$
$$S_2O_3^{2-} + 5\text{H}_2\text{O} \rightarrow 2\text{SO}_4^{2-} + 10\text{H}^+ + 8e^-$$

Write the overall equation for the reaction.

(1)

(d) The boiling temperatures of the hydrogen halides are shown.

<table>
<thead>
<tr>
<th>Hydrogen halide</th>
<th>Boiling temperature / K</th>
</tr>
</thead>
<tbody>
<tr>
<td>HF</td>
<td>293</td>
</tr>
<tr>
<td>HCl</td>
<td>188</td>
</tr>
<tr>
<td>HBr</td>
<td>206</td>
</tr>
<tr>
<td>HI</td>
<td>238</td>
</tr>
</tbody>
</table>

*(i)* London forces are present in all of these compounds.

Describe how these forces arise.

(2)
(ii) State why the London forces are greater in hydrogen iodide than in hydrogen bromide.

.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...

(1)

(iii) Explain why the boiling temperature of hydrogen fluoride is higher than that of hydrogen chloride.

.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...

(2)

(e) In the solid state, phosphorus(V) chloride exists as [PCl₄]⁺ and [PCl₆]⁻ ions.

Predict the shapes of these ions. Fully justify your answers.

Shape [PCl₄]⁺ .................................................................
Shape [PCl₆]⁻ .................................................................

Justification

.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...
.......................................................................................................................... ...

(Total for Question 22 = 14 marks)
23 This question is about mechanisms involving halogenoalkanes.

(a) Bromoethane reacts with dilute aqueous potassium hydroxide in a nucleophilic substitution reaction to form ethanol.

(i) Complete the mechanism for the reaction by adding curly arrows and the relevant dipole.

![Mechanism Diagram]

(ii) Explain the meaning of the term **nucleophilic substitution** in this mechanism.
(b) Chlorofluorocarbons, CFCs, were used for refrigerants, solvents and aerosol propellants because they are unreactive and neither flammable nor toxic.

However, in the stratosphere, ultraviolet radiation breaks CFCs into free radicals and these react with ozone.

Write the equation for the formation of two free radicals from a molecule of chlorotrifluoromethane, CF₃Cl. Curly arrows are not required.

\[ \text{(1)} \]

(Total for Question 23 = 6 marks)

TOTAL FOR SECTION B = 39 MARKS
SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

Many organic compounds have characteristic odours.
Some of these odours are pleasant, and the organic compounds are used in perfumes, soaps, deodorants, shampoos and other cosmetics.
Limonene is a colourless liquid which is present in the rind of lemons.

Linalool occurs in lavender oil.

Geraniol and citronellol occur in lemon grass.
They have rose-like odours.
(a) (i) Give the **molecular** formula for linalool.

(ii) Give the **empirical** formula for limonene.

(iii) Which of these four compounds are structural isomers?

(iv) Which of these four compounds show(s) geometric isomerism?

(b) Describe simple test tube reactions to identify the two functional groups present in linalool.

Give the reagents required and the observations you would make.
(c) (i) Explain whether it is possible to distinguish between limonene, linalool, geraniol and citronellol using **only** infrared spectroscopy.

(ii) Describe a chemical test that could be used to distinguish between samples of linalool and geraniol. Give the result of the test for both compounds.
(d) The four organic compounds react with hydrogen gas, in the presence of a suitable catalyst.

(i) Name a suitable catalyst for the reaction with hydrogen. 

(ii) Complete the balanced equation for the reaction of linalool with excess hydrogen. 

\[
\text{OH} \\
\begin{array}{c}
\text{C} \\
\text{C} \\
\text{C} \\
\text{C} \\
\end{array}
\]
(iii) A sample of lavender oil contained 70.0% by mass of linalool and no other unsaturated compounds. Calculate the minimum volume of hydrogen gas, measured at room temperature and pressure, needed to completely reduce 2.55 g of this lavender oil.

(The molar volume of hydrogen at room temperature and pressure is 24.0 dm$^3$ mol$^{-1}$. The molar mass of linalool is 154 g mol$^{-1}$)
(e) Hydrogen bromide reacts with $\text{C} \equiv \text{C}$ bonds such as those in citronellol.

Draw the mechanism for the reaction of hydrogen bromide with citronellol.

You should use the formula

\[
\begin{array}{c}
\text{CH}_3 \\
\text{C} \equiv \\ \\
\text{H}
\end{array}
\]

\[
\begin{array}{c}
\text{CH}_3 \\
\text{C} \equiv \\ \\
\text{R}
\end{array}
\]

\[
\begin{array}{c}
\text{CH}_3 \\
\text{C} \equiv \\ \\
\text{H}
\end{array}
\]

to represent a molecule of citronellol.

Include the dipole on the hydrogen bromide molecule.
The Periodic Table of Elements

<table>
<thead>
<tr>
<th>1</th>
<th>2</th>
<th>3</th>
<th>4</th>
<th>5</th>
<th>6</th>
<th>7</th>
<th>8</th>
</tr>
</thead>
<tbody>
<tr>
<td>H</td>
<td>1.0</td>
<td>1</td>
<td>He</td>
<td>4.0</td>
<td>18</td>
<td>Ne</td>
<td>10</td>
</tr>
<tr>
<td>Li</td>
<td>6.9</td>
<td>3</td>
<td>Be</td>
<td>9.0</td>
<td>4</td>
<td>B</td>
<td>10.8</td>
</tr>
<tr>
<td>Na</td>
<td>23.0</td>
<td>11</td>
<td>Mg</td>
<td>24.3</td>
<td>12</td>
<td>Al</td>
<td>27.0</td>
</tr>
<tr>
<td>K</td>
<td>39.1</td>
<td>19</td>
<td>Ca</td>
<td>40.1</td>
<td>20</td>
<td>Sc</td>
<td>45.0</td>
</tr>
<tr>
<td>Rb</td>
<td>85.5</td>
<td>37</td>
<td>Sr</td>
<td>87.6</td>
<td>36</td>
<td>Y</td>
<td>88.9</td>
</tr>
<tr>
<td>Cs</td>
<td>132.9</td>
<td>55</td>
<td>Ba</td>
<td>137.3</td>
<td>56</td>
<td>La*</td>
<td>138.9</td>
</tr>
<tr>
<td>Fr</td>
<td>[223]</td>
<td>87</td>
<td>Ra</td>
<td>[226]</td>
<td>88</td>
<td>Ac*</td>
<td>[227]</td>
</tr>
</tbody>
</table>

* Lanthanide series
* Actinide series

Elements with atomic numbers 112-116 have been reported but not fully authenticated.