

**Paper Reference(s) 9CH0/03**  
**Pearson Edexcel Level 3 GCE**

Total Marks
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**Chemistry**

**Advanced**

**PAPER 3: General and Practical Principles in Chemistry**

**Time: 2 hours 30 minutes**

**In the boxes below, write your name, centre number and candidate number.**

<b>Surname</b>					
<b>Other names</b>					
<b>Centre Number</b>					
<b>Candidate Number</b>					



**YOU MUST HAVE**

**Scientific calculator, Data Booklet, ruler**

**YOU WILL BE GIVEN**

**Diagram Booklet, Periodic Table**

**INSTRUCTIONS**

**Answer ALL questions.**

**Answer the questions in the spaces provided in this Question Paper or in the separate Diagram Booklet – there may be more space than you need.**

**INFORMATION**

**The total mark for this paper is 120.**

**The marks for EACH question are shown in brackets – use this as a guide as to how much time to spend on each question.**

**For the question marked with an ASTERISK (\*), marks will be awarded for your ability to structure your answer logically showing the points that you make are related or follow on from each other where appropriate.**

**A Periodic Table is provided as a separate insert.**

**There may be spare copies of some diagrams.**

**ADVICE**

**Read each question carefully before you start to answer it.**

**Show all your working in calculations and include units where appropriate.**

**Check your answers if you have time at the end.**

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**Answer ALL questions.**

**Write your answers in the spaces provided.**

**1 Relative atomic mass is an important concept in chemistry.**

**(a) Define the term relative atomic mass.  
(2 marks)**

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**(continued on the next page)**

**1 continued.**

**(b) Look at the table for Question 1(b) in the Diagram Booklet. A sample of neon consisted of three isotopes.**

**Calculate the relative atomic mass of neon in this sample.**

**Give your answer to three significant figures.  
(2 marks)**

**(Total for Question 1 = 4 marks)**

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**Turn over**

**2 Ammonium cobalt(II) sulfate is made by mixing aqueous solutions of ammonium sulfate and excess cobalt(II) sulfate.**

**(a) Dry crystals of ammonium cobalt(II) sulfate,  $(\text{NH}_4)_2\text{SO}_4 \cdot \text{CoSO}_4 \cdot 6\text{H}_2\text{O}$ , are obtained by the procedure shown.**

**Step 1** The reaction mixture is transferred to an evaporating basin, heated gently and then left to crystallise.

**Step 2** The crystals are separated by gravity filtration.

**Step 3** The crystals are then RINSED with a small amount of ICE-COLD water.

**Step 4** The rinsed crystals are placed in a WARM OVEN for 30 minutes.

**(continued on the next page)**



**2 continued.**

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**2 continued.**

**(b) The percentage yield of this reaction is 70.0%.**

**Give TWO possible reasons, other than an incomplete reaction, why the yield is less than 100%.**

**(2 marks)**

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**(Total for Question 2 = 12 marks)**

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**3 continued.**

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**(ii) The colour of the flame is yellow.**

**Give the FORMULA of the metal ion present  
in salt X.  
(1 mark)**

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3 continued.

- (b) A sample of **X** is placed in a test tube and dissolved in deionised water.  
The solution is acidified with hydrochloric acid and barium chloride solution is added.

A white precipitate forms.

- (i) Give the **FORMULA** of the anion present in **X**.  
(1 mark)

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- (ii) Deduce the **FORMULA** of **X**, using your answers to (a)(ii) and (b)(i).  
(1 mark)

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**3 continued.**

**(c) Y is identified as hydrated potassium carbonate,  $K_2CO_3 \cdot nH_2O$ .**

**Look at the diagram for Question 3(c) in the Diagram Booklet. Two of the students were asked to determine the number of moles of water of crystallisation,  $n$ , in Y using the procedure shown:**

- weigh a sample of hydrated Y into a pre-weighed crucible**
  - place a lid loosely on the crucible and heat it for five minutes to remove the water of crystallisation**
  - allow the crucible and lid to cool, remove the lid and then reweigh the crucible with its contents.**
- (i) The first student carried out the experiment but forgot to use the lid.**

**Explain how this mistake would affect the calculated value of  $n$ .**

**(2 marks)**

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**Turn over**

**3 continued.**

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**(ii) The second student carried out the experiment but heated the apparatus for only ONE minute.**

**Explain how this mistake would affect the calculated value of  $n$ .**

**(2 marks)**

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**3 continued.**

**(iii) In an accurate experiment, Y is found to consist of 71.9%  $\text{K}_2\text{CO}_3$  by mass.**

**Calculate the value of n.  
(3 marks)**

**(Total for Question 3 = 13 marks)**

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**4 This question is about the white solid barium carbonate.**

**Look at the equation and table for Question 4 in the Diagram Booklet.**

**(a) Barium carbonate decomposes under suitable conditions to form barium oxide and carbon dioxide.**

**Standard molar entropy data related to this reaction are shown.**

**(i) Show that barium carbonate is thermally stable at 298 K, using the data in the equation and in the table.  
(5 marks)**

**4 continued.**

**(continued on the next page)**

**4 continued.**

- (ii) Calculate the lowest temperature, in °C, at which it is thermodynamically feasible for barium carbonate to decompose. Give your answer to three significant figures. (3 marks)**

**(continued on the next page)**















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**(Total for Question 5 = 6 marks)**

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6 An ester **Q** has the molecular formula  $C_8H_{16}O_2$

(a) When burned in excess oxygen, 1.879 g of **Q** formed 4.594 g of carbon dioxide and 1.879 g of water.

Show that the empirical formula of **Q** is  $C_4H_8O$ .  
(4 marks)

**6 continued.**

**(continued on the next page)**

**6 continued.**

**(b) Look at the table for Question 6(b) in the Diagram Booklet. Data from the high resolution  $^1\text{H}$  (proton) NMR spectrum of the ester **Q** are shown in the table.**

**Part of the structure of **Q** is shown.**

**Look at the diagram for Question 6(b) in the Diagram Booklet. Complete the structure of **Q**.**

**Justify your answer by linking the proton environments in your structure to the relative peak areas and the splitting pattern of the peaks.  
(7 marks)**

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**(Total for Question 6 = 11 marks)**

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**7 This question is about some reactions of carbonyl compounds.**

**(a) Look at the diagram for Question 7(a) in the Diagram Booklet. Methyl methacrylate is the monomer used to make the polymer perspex. It can be synthesised from propanone using the reaction scheme shown.**

**(continued on the next page)**

**7 continued.**

- (i) Draw the mechanism for the reaction in Step 1.  
Include curly arrows and any relevant lone pairs and dipoles.  
(4 marks)**

**7 continued.**

**(ii) Look at the table for Question 7(a)(ii) in the Diagram Booklet. Complete the table to show the information missing from the reaction scheme.**

**(6 marks)**

**(iii) Look at the diagram for Question 7(a)(iii) in the Diagram Booklet. Complete the equation for the formation of the polymer from methyl methacrylate.**

**(2 marks)**

**(b) Propanone can be formed from the fermentation of polysaccharides such as starch. The propanone can be separated from the fermentation mixture by distillation.**

**Using the space provided in the Diagram Booklet for Question 7(b), draw the apparatus used in the laboratory for distillation of propanone from the reaction mixture.**

**(3 marks)**

**(continued on the next page)**

**7 continued.**

- (c) Carbonyl compounds, such as propanone, react with 2,4-dinitrophenylhydrazine in solution (Brady's reagent) to form a precipitate which can be used to identify the compound.**

**The precipitate can be purified by recrystallisation.**

**Details of the recrystallisation process are shown.**

- Step 1 Dissolve the precipitate in the minimum volume of hot ethanol.**
- Step 2 Warm a filter paper and funnel in an oven for use in Step 3.**
- Step 3 Filter the solution whilst still warm to remove any undissolved solids, using gravity filtration.**
- Step 4 Allow the filtrate to cool and recrystallise.**
- Step 5 Filter the crystals under reduced pressure.**
- Step 6 Rinse the crystals with a small amount of ice-cold ethanol.**
- Step 7 Dry the crystals between filter papers and leave in a desiccator.**

7 continued.

- (i) Explain why the filter paper and funnel are warmed in an oven before Step 3.  
(2 marks)

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8 This question is about acids and bases.

- (a) Devise an experiment to determine the acid dissociation constant,  $K_a$ , for a solution of ethanoic acid,  $\text{CH}_3\text{COOH}$ , of unknown concentration.

Assume you have access to a pH meter and a solution of sodium hydroxide of similar concentration to the acid.

Include how to determine  $K_a$  from your results.  
(5 marks)

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8 continued.

(b)  $500 \text{ cm}^3$  of a buffer solution of  $\text{pH} = 4.70$  is required.

Calculate the volume of  $0.800 \text{ mol dm}^{-3}$  sodium ethanoate solution and of  $0.800 \text{ mol dm}^{-3}$  ethanoic acid needed to make this buffer.  
(3 marks)

$[\text{K}_a \text{ for ethanoic acid} = 1.74 \times 10^{-5} \text{ mol dm}^{-3}]$

**8 continued.**

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8 continued.

- (c) Calculate the pH of the solution formed when  $51.2 \text{ cm}^3$  of  $0.927 \text{ mol dm}^{-3}$   $\text{NaOH}(\text{aq})$  is mixed with  $40.4 \text{ cm}^3$  of  $0.370 \text{ mol dm}^{-3}$   $\text{H}_2\text{SO}_4(\text{aq})$ .  
(6 marks)

[Ionic product of water  
 $K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ ]

**8 continued.**

**(Total for Question 8 = 14 marks)**

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9 Pineapple juice contains the weak acids citric acid ( $\text{C}_6\text{H}_8\text{O}_7$ ) and ascorbic acid ( $\text{C}_6\text{H}_8\text{O}_6$ ). The amount of each compound in a sample of  $150\text{ cm}^3$  of pineapple juice can be determined by titration.

(a) Experiment 1 is designed to determine the total amount of acid.

$10.0\text{ cm}^3$  samples of pineapple juice are transferred to separate conical flasks and titrated with a solution of sodium hydroxide of known concentration.

The total amount of acid in the  $150\text{ cm}^3$  sample of pineapple juice is  $8.00 \times 10^{-3}\text{ mol}$ .

(i) Give a reason why methyl orange would NOT be a suitable indicator to use in this titration.  
(1 mark)

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**9 continued.**

- (ii) A student did not notice an air bubble in the tip of the burette BEFORE carrying out one of their accurate titrations. During this titration, the air bubble escaped.**

**Explain the effect this mistake would have on the value of this titre.**

**(2 marks)**

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**9 continued.**

**(b) Experiment 2 is carried out to determine the amount of ascorbic acid ( $C_6H_8O_6$ ) in the pineapple juice.**

**An outline procedure for this experiment is given.**

**Step 1** 5.00 cm<sup>3</sup> of the pineapple juice is added to a conical flask.

**Step 2** Deionised water, a small amount of HCl(aq), a few crystals of potassium iodide, KI, and 3 drops of starch solution are also added to the flask.

**Step 3** The contents of the flask are swirled to ensure the KI dissolves fully.

**Step 4** The resultant mixture is titrated with a solution of potassium iodate(V),  $KIO_3$ (aq), of concentration 0.00100 mol dm<sup>-3</sup>.

**The reactions that take place are on page 17 in the Diagram Booklet.**

**Only the ascorbic acid reacts with the iodine.**

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**9 continued.**

**(ii) The TOTAL amount of acid in the  $150\text{ cm}^3$  sample is  $8.00 \times 10^{-3}\text{ mol}$ .**

**The mean titre in Experiment 2 using  $5.00\text{ cm}^3$  of pineapple juice is  $9.50\text{ cm}^3$ .**

**Calculate the mass of CITRIC ACID in the  $150\text{ cm}^3$  sample.**

**(5 marks)**

**9 continued.**

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**Turn over**

**9 continued.**

- (c) Look at the diagram for Question 9(c) in the Diagram Booklet. While doing background research for the experiment, a student found that three other compounds, D, E and F, are often present in pineapple juice.**

**Predict which one of these compounds is most likely to affect the result of Experiment 1 and hence predict the effect on the mass of citric acid calculated in (b)(ii).**

**Justify your answer.**

**(3 marks)**

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**(Total for Question 9 = 14 marks)**

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**10** The progress of the reaction between iodine and propanone with an acid catalyst can be followed in an experiment using a titrimetric method.

### **Procedure**

- Step 1** Mix  $25\text{ cm}^3$  of  $1\text{ mol dm}^{-3}$  aqueous propanone with  $25\text{ cm}^3$  of  $1\text{ mol dm}^{-3}$  sulfuric acid in a beaker. Both these reactants are in excess.
- Step 2** Start the stop clock as  $50\text{ cm}^3$  of  $0.02\text{ mol dm}^{-3}$  iodine solution is added to the beaker. Mix the reactants thoroughly.
- Step 3** Withdraw a  $10.0\text{ cm}^3$  sample of the reaction mixture, using a pipette, and transfer it to a conical flask.
- Step 4** Add a spatula measure of sodium hydrogencarbonate, noting the exact time.
- Step 5** Titrate the iodine present in the  $10.0\text{ cm}^3$  sample with  $0.01\text{ mol dm}^{-3}$  sodium thiosulfate solution, using starch indicator.
- Step 6** Continue to withdraw  $10.0\text{ cm}^3$  samples about every two minutes, repeating Steps 4 and 5 with each sample.

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**10 continued.**

- (ii) Write the IONIC equation for the reaction that takes place during Step 4. State symbols are not required.  
(1 mark)**

**(continued on the next page)**

10 continued.

(b) Some data from the experiment are shown.

Time sodium hydrogencarbonate is added / min	2.0	5.0	6.5	8.0	10.5	12.0
Volume of sodium thiosulfate / cm <sup>3</sup>	19.2	15.5	14.0	12.1	9.5	7.2

- (i) Look at the grid for Question 10(b)(i) in the Diagram Booklet. Plot a graph of the volume of sodium thiosulfate against the time the sodium hydrogencarbonate is added.  
(2 marks)
- (ii) Explain how the graph of volume of thiosulfate against time confirms the reaction is zero order with respect to iodine, I<sub>2</sub>.  
(3 marks)

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**10 continued.**

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10 continued.

- (c) The overall rate equation for the reaction is  
 $\text{rate} = k[\text{H}^+(\text{aq})][\text{CH}_3\text{COCH}_3(\text{aq})]$ .

Look at the diagram for Question 10(c) in the Diagram Booklet. A student researching the mechanism for the reaction found this example.

- (i) Predict which of the three steps is the rate-determining step.  
Justify your answer.  
(2 marks)

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