

Write your name here

Surname

Other names

Centre Number

Candidate Number

**Edexcel GCE**

# Chemistry

**Advanced**

**Unit 4: General Principles of Chemistry I – Rates,  
Equilibria and Further Organic Chemistry  
(including synoptic assessment)**

Thursday 26 January 2012 – Afternoon

**Time: 1 hour 40 minutes**

Paper Reference

**6CH04/01**

**You must have: Data Booklet**

Total Marks

**Candidates may use a calculator.**

## Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*

## Information

- The total mark for this paper is 90.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (\*) are ones where the quality of your written communication will be assessed – *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*
- A Periodic Table is printed on the back cover of this paper.

## Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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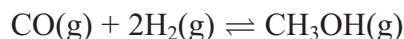


**PEARSON**

## SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 20 minutes on this section. For each question, select one answer from A to D and put a cross in the box . If you change your mind, put a line through the box  and then mark your new answer with a cross .

- 1 The reaction between carbon monoxide and hydrogen reaches a dynamic equilibrium.



- (a) Which of these statements about a dynamic equilibrium is **not** true?

(1)

- A The forward rate of reaction is equal to the backward rate of reaction.
- B The concentrations of the products and reactants do not change.
- C The concentrations of the products and reactants are equal.
- D The equilibrium can be approached from either direction.

- (b) The  $K_c$  expression for the above reaction is

(1)

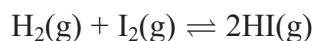
- A  $K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}] \times [\text{H}_2]^2}$
- B  $K_c = \frac{[\text{CO}] \times 2[\text{H}_2]}{[\text{CH}_3\text{OH}]}$
- C  $K_c = \frac{[\text{CO}] \times [\text{H}_2]^2}{[\text{CH}_3\text{OH}]}$
- D  $K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}] \times 2[\text{H}_2]}$

(Total for Question 1 = 2 marks)

Use this space for any rough working. Anything you write in this space will gain no credit.



- 2 Hydrogen and iodine, both with an initial concentration of  $0.010 \text{ mol dm}^{-3}$ , were allowed to react. At equilibrium, the concentration of hydrogen iodide was  $0.0030 \text{ mol dm}^{-3}$ .

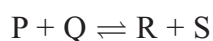


$K_c$  is calculated using the values

		$\text{H}_2(\text{g}) / \text{mol dm}^{-3}$	$\text{I}_2(\text{g}) / \text{mol dm}^{-3}$	$\text{HI}(\text{g}) / \text{mol dm}^{-3}$
<input type="checkbox"/>	A	0.0070	0.0070	0.0030
<input type="checkbox"/>	B	0.0040	0.0040	0.0030
<input type="checkbox"/>	C	0.0040	0.0040	0.0060
<input type="checkbox"/>	D	0.0085	0.0085	0.0030

(Total for Question 2 = 1 mark)

- 3 The reaction below reached a dynamic equilibrium from an initial mixture of all four substances P, Q, R and S in aqueous solution.



The following data were obtained.

Substance	Concentration at equilibrium / $\text{mol dm}^{-3}$
P	0.050
Q	0.040
R	0.020
S	0.010

$K_c$  for the equilibrium is

- A 0.10
- B 0.33
- C 3.00
- D 10.0

(Total for Question 3 = 1 mark)



4 The Haber process is used to make ammonia from nitrogen and hydrogen at 450 °C.



(a) If the partial pressures of these gases were measured in atm, the units of the equilibrium constant  $K_p$  will be

(1)

- A atm
- B  $\text{atm}^2$
- C  $\text{atm}^{-2}$
- D  $\text{atm}^{-1}$

(b) When the temperature of the system is increased

(1)

- A  $K_p$  decreases.
- B  $K_p$  increases.
- C  $K_p$  stays the same.
- D  $K_p$  first decreases and then increases.

(Total for Question 4 = 2 marks)

5 In high performance liquid chromatography, HPLC, which of these factors does **not** affect the time taken for a component to pass through the column?

- A Type of detector
- B Material of stationary phase
- C Particle size of stationary phase
- D Temperature of column

(Total for Question 5 = 1 mark)

6 When equimolar amounts of the solutions below are mixed, which forms a buffer solution with a pH less than 7?

- A Hydrochloric acid and sodium chloride
- B Ethanoic acid and sodium ethanoate
- C Sodium hydroxide and sodium chloride
- D Ammonia and ammonium chloride

(Total for Question 6 = 1 mark)



7 The pH of a  $1.5 \text{ mol dm}^{-3}$  solution of hydrochloric acid,  $\text{HCl(aq)}$ , is

- A -1.50
- B -0.18
- C 0.18
- D 1.50

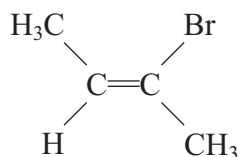
(Total for Question 7 = 1 mark)

8 Which of these solid substances is likely to have the greatest standard entropy? Use of the data booklet is not required.

- A  $\text{SnO}$
- B  $\text{SnO}_2$
- C  $\text{SnBr}_2$
- D  $\text{SnBr}_4$

(Total for Question 8 = 1 mark)

9 What is the correct name for the molecule shown below?



- A *Z*-2-bromobut-2-ene
- B *E*-2-bromobut-2-ene
- C *E*-3-bromobut-2-ene
- D *Z*-3-bromobut-2-ene

(Total for Question 9 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.



10 Ketones react with hydrogen cyanide, HCN, in the presence of cyanide ions, CN<sup>-</sup>.

(a) Which of these ketones does **not** form a racemic mixture in this reaction?

(1)

- A CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>COCH<sub>3</sub>
- B CH<sub>3</sub>CH<sub>2</sub>COCH<sub>2</sub>CH<sub>3</sub>
- C CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>COCH<sub>3</sub>
- D CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>COCH<sub>2</sub>CH<sub>3</sub>

(b) This type of reaction is classified as

(1)

- A nucleophilic substitution.
- B nucleophilic addition.
- C electrophilic addition.
- D electrophilic substitution.

(Total for Question 10 = 2 marks)

11 Which of these is **not** observed when ethanoyl chloride reacts with water?

- A Misty fumes given off.
- B The gas given off turns damp blue litmus paper red.
- C The mixture gets hot.
- D A white precipitate forms.

(Total for Question 11 = 1 mark)

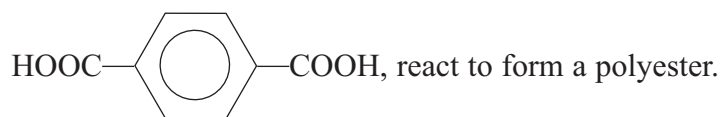
12 UV light is useful in initiating some reactions because it

- A lowers the activation energy of the reaction.
- B causes bonds in molecules to stretch and bend.
- C causes molecules to form ions.
- D causes molecules to form free radicals.

(Total for Question 12 = 1 mark)

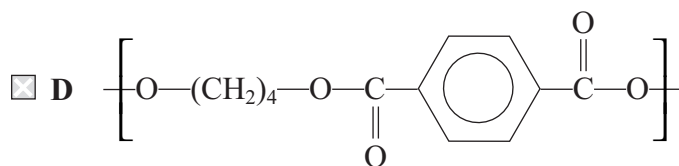
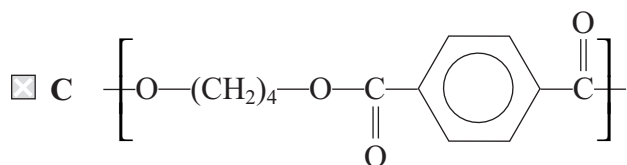
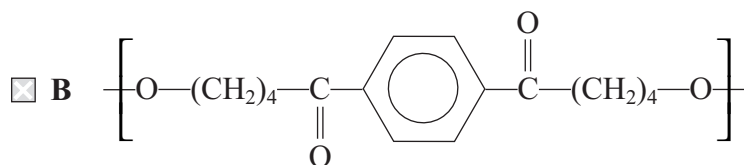
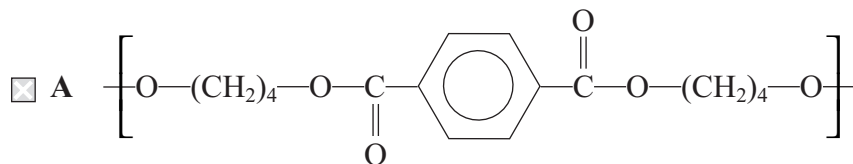


13 Butane-1,4-diol,  $\text{HO}(\text{CH}_2)_4\text{OH}$ , and benzene-1,4-dicarboxylic acid,



(a) The repeat unit of the polyester is

(1)



(b) The type of reaction is

(1)

- A hydrolysis.
- B addition.
- C substitution.
- D condensation.

(Total for Question 13 = 2 marks)

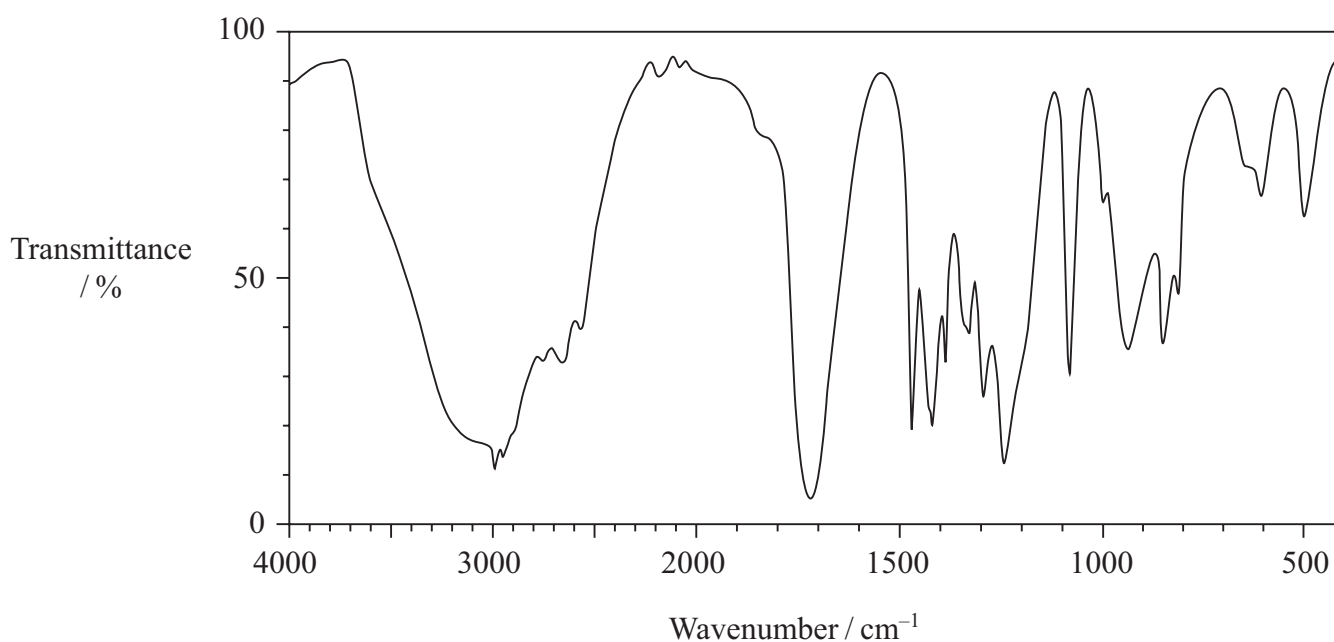


14 The equation for the enthalpy of hydration for a magnesium ion is

- A  $\text{Mg}^{2+}(\text{s}) + \text{aq} \rightarrow \text{Mg}^{2+}(\text{aq})$
- B  $\text{Mg}^{2+}(\text{g}) + \text{aq} \rightarrow \text{Mg}^{2+}(\text{aq})$
- C  $\text{Mg}^{2+}(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{g}) + \text{aq}$
- D  $\text{Mg}^{2+}(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{s}) + \text{aq}$

(Total for Question 14 = 1 mark)

15 The IR spectrum of a substance is shown below.



Which of the following substances has this spectrum?

You may use the information on page 6 of the data booklet.

- A Propan-1-ol
- B Propanal
- C Propanone
- D Propanoic acid

(Total for Question 15 = 1 mark)





16 Two ketones,  $\text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_3$  and  $\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3$ , both have  $M_r = 86$ . Which peak due to fragmentation into singly charged ions would you expect to be present in the mass spectrum of one but not the other?

- A 71
- B 57
- C 43
- D 29

(Total for Question 16 = 1 mark)

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**TOTAL FOR SECTION A = 20 MARKS**



## SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

17 Two organic compounds, X and Y, both with the molecular formula  $C_4H_8O$ , contain a carbonyl group.

(a) Describe what you would see when 2,4-dinitrophenylhydrazine is added to either of these compounds.

(1)

(b) It is suspected that X is a ketone and Y is an aldehyde. Outline a chemical test you could carry out to confirm this, describing the results in each case.

(3)

(c) (i) Give the structural formulae of the two possible isomers of Y which are aldehydes.

(1)

(ii) Name the technique you would use to purify the product of the test with 2,4-dinitrophenylhydrazine.

(1)

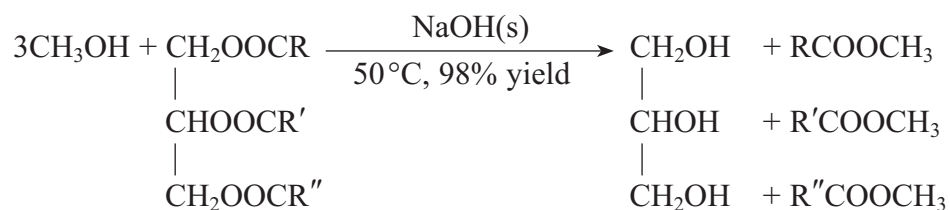
(iii) Other than by spectroscopic techniques, how would you use the purified product to identify compound Y? [Practical details are not required.]

(2)

(Total for Question 17 = 8 marks)



18 Kits for manufacturing biodiesel from vegetable oils and methanol are sold for home use. The reaction which takes place may be represented by the following equation.



\*(a) Describe any two of the main hazards when carrying out this reaction. What precaution would you take to minimise the risk in each case?

(4)

Hazard .....

Precaution .....

Hazard .....

Precaution .....

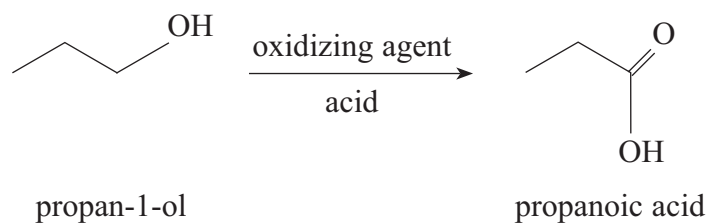
(b) Suggest **two** environmental benefits of using these kits, despite the associated risks.

(2)

(Total for Question 18 = 6 marks)



19 The carboxylic acid, propanoic acid, can be prepared by oxidation of the alcohol, propan-1-ol.



(a) (i) Identify a suitable oxidizing agent you could use in this reaction.

(1)

(ii) If you carried out this preparation in the laboratory, describe **two** measures you would take to ensure the maximum possible yield of propanoic acid is obtained.

(2)

(iii) Propanoic acid can be made by the hydrolysis of a nitrile. Give the structural formula of the nitrile and write an equation for this reaction.

(3)

Structural formula .....

Equation



\*(b) Propanoic acid reacts with methanol, CH<sub>3</sub>OH, to form the ester, methyl propanoate.



Even with the use of a catalyst, this reaction is quite slow and incomplete. Suggest a reagent, to replace the propanoic acid, which would form the ester at a faster rate. Suggest **two** reasons why your chosen reagent reacts faster.

(3)

(c) The structure of methyl propanoate can be investigated by using high resolution <sup>1</sup>H nuclear magnetic resonance (nmr) spectroscopy.

(i) What type of radiation interacts with <sup>1</sup>H nuclei in nmr spectroscopy?

(1)

(ii) Describe what happens to <sup>1</sup>H nuclei when they absorb this radiation.

(2)

(iii) Complete the table to show values for the chemical shift of the different <sup>1</sup>H nuclei in methyl propanoate and their splitting pattern. Page 7 of the data booklet gives information about chemical shifts.

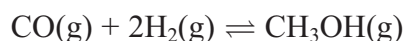
(2)

<sup>1</sup> H environment	Chemical shift, δ / ppm	Splitting pattern
CH <sub>3</sub> O-	3.7	Singlet
-CH <sub>2</sub> -	2.3	
-CH <sub>3</sub>		Triplet

(Total for Question 19 = 14 marks)



20 The exothermic reaction between carbon monoxide and hydrogen can be used industrially to make methanol. The process is carried out at 250 °C and between 50 and 100 atm.



(a) Explain why increasing the pressure increases the yield of methanol. Give **one** disadvantage of increasing the pressure.

(2)

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(b) The reaction gives a greater equilibrium yield at 100 °C than at 250 °C.

(i) Explain, in terms of the entropy change of the surroundings and the total entropy change of the reaction, why this is so.

A calculation is **not** required.

(2)

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(ii) Explain why the reaction is, nevertheless, carried out at 250 °C.

(1)

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(c) Given that the reaction is an equilibrium, suggest **two** ways in which the atom economy of this process could be maximised without changing the temperature or pressure.

(2)

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(Total for Question 20 = 7 marks)



21 This question is about the kinetics of the reaction between bromoethane and aqueous hydroxide ions.

(a) The results of an experiment to find the initial rate of the reaction are shown in the table below.

$[\text{CH}_3\text{CH}_2\text{Br}]$ / $\text{mol dm}^{-3}$	$[\text{OH}^-]$ / $\text{mol dm}^{-3}$	Initial rate / $\text{mol dm}^{-3} \text{ s}^{-1}$
0.100	0.150	$1.54 \times 10^{-6}$

The rate equation for the reaction is

$$\text{rate} = k[\text{CH}_3\text{CH}_2\text{Br}][\text{OH}^-]$$

(i) Calculate the value of  $k$ . Give your answer to three significant figures and include units.

(3)

(ii) Calculate the initial rate if the concentrations of both reactants were changed to  $0.020 \text{ mol dm}^{-3}$ .

(1)

(b) (i) State the order of the reaction.

(1)

(ii) The mechanism for this reaction can be inferred from the rate equation. Draw the transition state formed in the reaction between bromoethane and hydroxide ions.

(2)



(c) The rate constant for the reaction between bromoethane and hydroxide ions was determined at five different temperatures. The results are shown in the table below.

Temperature (T) / K	1/Temperature (1/T) / K <sup>-1</sup>	Rate constant, <i>k</i>	ln <i>k</i>
293	$3.41 \times 10^{-3}$	$5.83 \times 10^{-5}$	-9.75
303	$3.30 \times 10^{-3}$	$1.67 \times 10^{-4}$	-8.70
313	$3.19 \times 10^{-3}$	$5.26 \times 10^{-4}$	-7.55
323	$3.10 \times 10^{-3}$	$1.36 \times 10^{-3}$	-6.60
333		$3.77 \times 10^{-3}$	

(i) Complete the missing values in the table.

(2)

(ii) Plot a graph of ln *k* against 1/*T*. Calculate the gradient of your graph and use this to calculate the activation energy, *E*<sub>A</sub>. The Arrhenius equation can be expressed as

$$\ln k = \frac{-E_A}{R} \times \left( \frac{1}{T} \right) + \text{a constant}$$

[Gas constant, *R* = 8.31 J K<sup>-1</sup> mol<sup>-1</sup>]

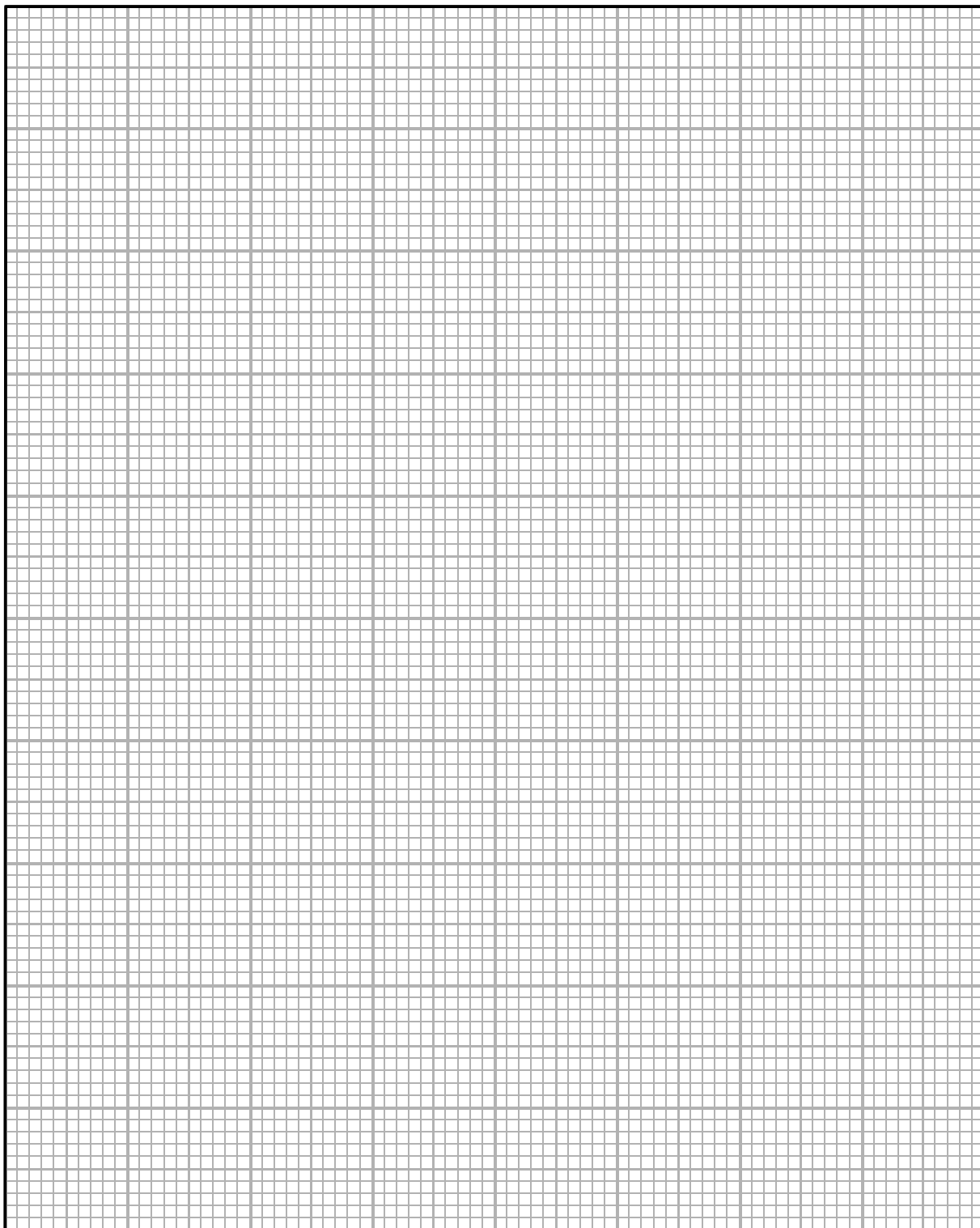
(5)





$1/T/K^{-1}$

$\ln k$



(Total for Question 21 = 14 marks)

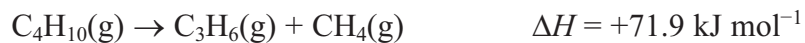
**TOTAL FOR SECTION B = 49 MARKS**



## SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

- 22 The hydrocarbon butane can be cracked to form propene and methane by passing it over a heated aluminium oxide catalyst at a temperature of 700 K. The equation for the reaction is



- (a) (i) Use page 20 of the data booklet to complete the table below.

(1)

Hydrocarbon	$S^\ominus / \text{J mol}^{-1} \text{K}^{-1}$
$\text{C}_4\text{H}_{10}(\text{g})$	+310.1
$\text{C}_3\text{H}_6(\text{g})$	+266.9
$\text{CH}_4(\text{g})$	

- (ii) Calculate the standard entropy change of the system,  $\Delta S_{\text{system}}^\ominus$ , for this reaction. Include a sign in your answer.

(2)

- (iii) Was the sign for your answer as you expected? Fully justify your answer.

(2)

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(iv) Calculate the entropy change of the surroundings,  $\Delta S_{\text{surroundings}}$ , at 700 K.

Include a sign and units in your answer.

Use this value and your answer to (ii) to explain why butane cracks into propene and methane at this temperature.

(3)

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(v) Calculate the minimum temperature needed for this reaction to be thermodynamically feasible.

(3)



(b) The aluminium oxide behaves as a heterogeneous catalyst. Explain both what is meant by the term **heterogeneous** and how, in terms of activation energy, the catalyst is able to speed up the reaction.

(3)

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**(Total for Question 22 = 14 marks)**



**23** The bubble bath ‘Colour Change Matey’ has amongst its ingredients the weak acid benzoic acid, as well as the indicator bromocresol green. When it is added to bath water, its colour changes from yellow to blue.

(a) (i) Write the  $K_a$  expression for the dissociation of benzoic acid,  $C_6H_5COOH$ . (1)

(ii) Use the data on page 18 of the data booklet to calculate the pH of a solution of benzoic acid,  $C_6H_5COOH$ , of concentration  $0.0025 \text{ mol dm}^{-3}$ . (2)

\*(b) Use the data on page 19 of the data booklet, and your answer to (a)(ii), to suggest why the bubble bath changes colour when it is diluted by being added to the bath water. (4)

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**(Total for Question 23 = 7 marks)**

**TOTAL FOR SECTION C = 21 MARKS**  
**TOTAL FOR PAPER = 90 MARKS**



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