IGCSE CHEMISTRY 4335, NOVEMBER 2005 MARK SCHEME

Paper 1F

1.	(a)	nucleus		**1**
	(b)	proton		**1**
	(c)	electron		**1**
	(d)	2		**1**
	(e)	helium		**1**
				Total 5 marks
2.	(a)	X		**1**
	(b)	20 / 21		**1**
	(c)	glowing splint		**1**
		relights		**1**
	(d)	carbon dioxide		**1**
				Total 5 marks
3.	(a)	zinc + sulphuric acid → zinc sulphate + hydrogen		**1**
	(b)	effervescence		**1**
		slower		**1**
		exothermic		**1**
	(c)	burning splint / flame		**1**
		squeaky pop / explosion		**1**
	(d)	(i) filtration		**1**
		(ii) barium chloride (solution)		**1**
		white precipitate		**1**
				Total 9 marks
4.	(a)	(i) coke		**1**
		limestone		**1**
		(ii) slag / calcium silicate		**1**
		(iii) $C(s) + O_2(g) \rightarrow CO_2(g)$		**1**
		formulae		**1**
		state symbols		**1**
		(iv) CO_2		**1**

(b) (i) air / oxygen **1**
water / moisture **1**
(ii) zinc **1**
prevents air/oxygen/water from contacting iron OR more reactive than iron **1**
(iii) (covered in) oil **1**

Total 11 marks

5. (a) bright/brilliant/blinding flame **1**
white solid/ ash/ smoke **1**

(b) water / H_2O **1**

(c) 11 **1**
solution/magnesium hydroxide is (weak) alkali / contains hydroxide ions **1**

(d) (i) hydrochloric (acid) **1**
(ii) neutralisation **1**
(iii) water **1**

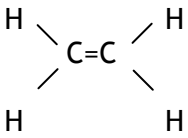
Total 8 marks

6. (a) hydrogen and carbon only **1**

(b) (i) alkanes **1**
(ii) C_nH_{2n+2} **1**
(iii) A and D / CH_4 and C_3H_8 **1**

(c) (i) compounds with same molecular formula **1**
but different structures/structural formulae **1**
(ii) none **1**

(d)



all atoms and bonds correct **1**
bond angles around C approximately 120° **1**

(e) 28 **1**

Total 10 marks

7. (a) nitrogen **1**
Air **1**
hydrogen **1**
natural gas / methane / hydrocarbons **1**
(P and Q can be in reverse order)
(b) (i) cooled **NOT** condensed **1**
(ii) liquid **1**

(c) (unused/recycled) nitrogen and hydrogen **1**

Total 7 marks

8. (a)

Name of substance	Ionic bonding	Covalent bonding	Insoluble in water	Soluble in water
ammonia		✓		✓
methane		✓		
poly(ethene)			✓	
sodium chloride	✓			✓
sodium hydroxide				

All six correct – 4 marks
5 or 4 correct – 3 marks
3 correct – 2 marks
2 correct – 1 mark **4**

(b) (i) any suitable use e.g. making bags/food packaging… **1**
(ii) any two from: soap, paper, ceramics, bleach, detergents **2**

Total 7 marks

9. (a) potassium manganate(VII) / manganese(IV) oxide **1**

(b) damp litmus paper **1**
bleached **1**

(c) (i) iron(III) chloride **1**
(ii) brown solid / precipitate **1**

(d) (i) iodine **1**
(ii) chlorine is more reactive (than iodine) **1**

Total 7 marks

10. (a) a shared pair of electrons **1**

(b) simple **1**
weak **1**
molecules **1**
low **1**

(c) (i) hydrogen shown with 1 electron **1**
oxygen shown as 2,6 **1**
(ii) one oxygen atom with two hydrogens **1**
each has full outer shell of electrons **1**
(iii) bent / v-shaped **1**

Total 10 marks

11. (a) electrons from Mg to F **1**
Mg loses 2 electrons **1**
each of two F gains 1 electron **1**

(b) Mg **1**
it has lost electrons **1**

(c) (i) Na^+F^- **1**
(ii) NaF **1**

(d) orange / yellow **1**

Total 8 marks

12. (a) (i) 5 **1**
(ii) colourless **1**

(b) (i) $NH_3 + HCl \rightarrow NH_4Cl$ **OR** $NH_4OH + HCl \rightarrow NH_4Cl + H_2O$ **2**

reagents **(1)** products **(1)**; **(-1)** for incorrect balancing.
(ii) (heat with) sodium hydroxide solution **1**
ammonia /alkaline gas given off **1**
test gas with damp U I / litmus paper – turns blue **1**
(iii) mix together same volumes **1**
no indicator/partial evaporation - not to dryness **1**
crystallise solution **1**
(**OR** if use indicator:
add charcoal
filter
evaporate/crystallise)

(c) (i) any soluble lead(II) salt **1**
any soluble chloride **1**
(ii) any equation that is cq on answer to c(i) **1**

Total 13 marks

Paper 2H

1. (a)

Name of substance	Ionic bonding	Covalent bonding	Insoluble in water	Soluble in water
ammonia		✓		✓
methane		✓		
poly(ethene)			✓	
sodium chloride	✓			✓
sodium hydroxide				

All six correct – 4 marks
5 or 4 correct – 3 marks
3 correct – 2 marks
2 correct – 1 mark **4**

(b) (i) any suitable use e.g. making bags/food packaging… **1**
(ii) any two from: soap, paper, ceramics, bleach, detergents **2**

Total 7 marks

2. (a) potassium manganate(VII) / manganese(IV) oxide **1**

(b) damp litmus paper **1**
bleached **1**

(c) (i) iron(III) chloride **1**
(ii) brown solid / precipitate **1**

(d) (i) iodine **1**
(ii) chlorine is more reactive (than iodine) **1**

Total 7 marks

3. (a) a shared pair of electrons **1**

(b) simple **1**
weak **1**
molecules **1**
low **1**

(c) (i) hydrogen shown with 1 electron **1**
oxygen shown as 2,6 **1**
(ii) one oxygen atom with two hydrogens **1**
each has full outer shell of electrons **1**
(iii) bent / v-shaped **1**

Total 10 marks

4. (a) electrons from Mg to F **1**
Mg loses 2 electrons **1**
each of two F gains 1 electron **1**

(b) Mg **1**
it has lost electrons **1**

(c) (i) Na^+F^- **1**
(ii) NaF **1**

(d) orange / yellow **1**

Total 8 marks

5. (a) (i) 5 **1**
(ii) colourless **1**

(b) (i) $NH_3 + HCl \rightarrow NH_4Cl$ **OR** $NH_4OH + HCl \rightarrow NH_4Cl + H_2O$ **2**
reagents **(1)** products **(1)**; **(-1)** for incorrect balancing.
(ii) (heat with) sodium hydroxide solution **1**
ammonia /alkaline gas given off **1**
test gas with damp U I / litmus paper – turns blue **1**
(iii) mix together same volumes **1**
no indicator/partial evaporation - not to dryness **1**
crystallise solution **1**
(**OR** if use indicator: add charcoal
filter
evaporate/crystallise)

(c) (i) any soluble lead(II) salt **1**
any soluble chloride **1**
(ii) any equation that is cq on answer to c(i) **1**

Total 13 marks

6. (a) NaCl(s) If state symbols missing / incorrect, **1**
$H_2O(l)$ **(1)** for both formulae correct **1**

(b) (i) silver nitrate (solution) **1**
(dilute) nitric acid **1**
(ii) white precipitate **1**
(iii) diffusion **1**

(c) (i) all three pieces drawn in correct sequence **1**
condenser at correct angle and connected via sidearm to rb flask with bung in neck of flask **1**
(**ALLOW** bung + thermometer in top of flask)
labels for sea water, cooling water and drinking water **1**
(ii) distillation **NOT** fractional distillation **1**

Total 10 marks

7. (a) Giant structure of (positive/metal/copper) ions **1**
electrons **1**
delocalised / free / mobile **1**

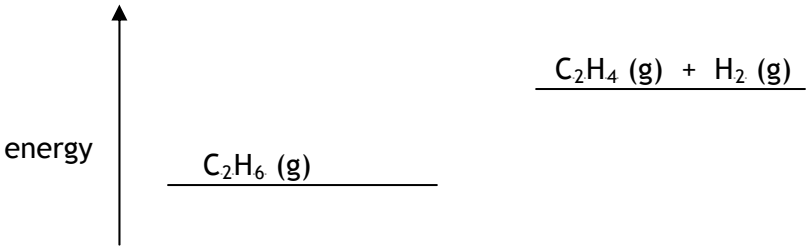
(b) (i) green **1**
black **1**
(ii) $CuCO_3 \rightarrow CuO + CO_2$ **1**
(iii) (bubble through) limewater **1**
turns milky/cloudy / white precipitate **1**
(iv) (dilute) nitric acid **1**
neutralisation **1**
(v) (pale) blue precipitate **1**
(vi) (dark) blue **1**
(vii) $[Cu(H_2O)_2(NH_3)_4]^{2+}$ **1**

(c) copper(I) oxide **1**
Cu_2O **1**

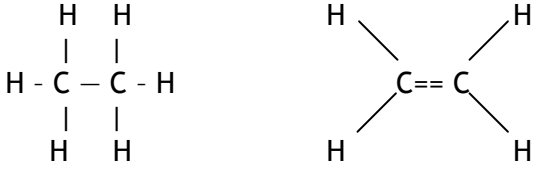
Total 15 marks

8. (a) (manufacture of) polymers / poly(ethene) / ethanol **1**
(manufacture of) ammonia / margarine / rocket fuel **1**

(b) **1**



(c) (i) **2**



(ii) bonds broken = 348 + (2×412) / 1172 **1**
bonds formed = 612 + 436 / 1048 **1**
energy change = 124 (kJ/mol) **1**

(d) increase in temperature
add catalyst
increase pressure
} any two for 1 each **2**

(e) (i) (⇌) reversible reaction **1**
(ii) (ΔH) enthalpy change / energy change / heat change **1**
increased **1**
decreased **1**

Total 14 marks

9. (a) fractional distillation **1**

(b) gasoline
kerosene
diesel
fuel oil
bitumen
} any two for 1 each **2**

(c) heat / high temperature / 200 – 400°C **1**
phosphoric acid **1**

(d) (i) sugar (cane) **1**
(ii) no crude oil **1**
plenty of land/suitable climate to grow sugar cane **1**

(e) (i) ethanol **1**
sulphuric/phosphoric/hydrochloric acid **1**
(ii) esters **1**

Total 11 marks

10. (a) effervescence / fizzing / bubbles
water goes cloudy / white precipitate
gets warmer
} any two for 1 each **2**
$Ca + 2H_2O \rightarrow Ca(OH)_2 + H_2$ **1**

(b) zinc oxide **1**
$Zn + H_2O \rightarrow ZnO + H_2$ **1**

(c) (i) $Zn + Fe^{2+} \rightarrow Zn^{2+} + Fe$ **Ignore state symbols** **1**
(ii) displacement / redox **1**

(d) oxygen / air **1**

(e) (i) (coated with) zinc **1**
(ii) zinc more reactive than iron **1**
zinc reacts/corrodes instead of iron **1**

Total 11 marks

11. (a) 160 **1**

(b) (i) $320000 \div 160$ **1**
$= 2000$ **1**
(ii) 2000×2 **1**
$= 4000$ **1**
(iii) 4000×56 **1**
$= 224000$ g $= 224$ (kg) **1**

(c) (i) it reduces the capacity of blood to carry oxygen / correct reference to haemoglobin **1**
(ii) $5000 \times 24 = 120000$ (dm^3) **1**

(d) $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ **2**
All formulae correct = 1, correct balancing = 1

(e) (i) silica / silicon dioxide / sand **1**

(ii) $CaCO_3 \rightarrow CaO + CO_2$ **1**

$CaO + SiO_2 \rightarrow CaSiO_3$ **1**

Total 14 marks

Paper 3

1. (a)

	Name of apparatus	Volume of liquid
A	measuring cylinder	8.6
B	beaker	100
C	pipette	25
D	burette	12.3

8

(b) (i) B **1**
(ii) D (burette) **1**

Total 10 marks

2. (a) 800
600
400
200
0
all 5 = 2 marks
any 3 = 1 mark **2**

(b) suitable y scale **1**
all points correct (-1 per error) **2**
best fit line attempted **1**

(c) graph turns
get better idea where maximum is **2**

(d) repeat **1**
check consistency of data **1**

(e) (i) no oxygen / hydrogen can not react / nothing to react with **1**
(ii) reference to 2:1 ratio in equation **2**
maximum amount of hydrogen reacting
(iii) 333 (accept 330 or 300) (cm^3) **1**

Total 14 marks

3. (a) when acid added **1**

(b) (i) circled anomalous point (at 5° C) **1**
(ii) got warmer **3**
absorbed heat/energy from surroundings
reaction faster
OR
too great a volume of solution(s)
bigger depth to look through
harder to see cross / less precipitate needed to obscure cross.
(iii) 7 seconds (from graph) **1**

(c) 1 / b(iii) cq marking **2**
correct answer (0.143 s^{-1})

(d) faster so percentage error in times greater **2**
loss of heat to surroundings

(e) (i) the higher the temp the faster the reaction **2**
idea of non linear

(ii) particles move faster **3**
more collisions per second / more frequent collisions
more collisions are successful / have energy > E_A

(f) volume thiosulphate constant
different volumes acid
different volumes water (to keep total volume / depth constant)
temp constant **4**

Total 19 marks

4. (a) 10.6 (g) **1**

(b) decreases **1**

(c) fume cupboard / well ventilated room **2**
sulphur dioxide is toxic.

(d) evaporation (of water) **1**
evaporation faster **1**

(e) measure pH **1**

Total 7 marks